

ELECTROCHEMISTRY (LECTURE-2)

Conductors $\Rightarrow$ Substances which allow the flow of current through them are called conductors. -
It is of two types:-

(i) Metallic conductors or Electronic conductors

- $\rightarrow$ Substances which conduct electricity due to ~~presence~~^{movement} of free electrons are called metallic conductors or electronic conductors.
- $\rightarrow$ conductivity decreases with rise in temperature due to increase resistance.
- $\rightarrow$ conductivity of metals are generally high.
- $\rightarrow$ They obey ohm's law but Faraday law of electricity do not obey.

(2) Electrolytic conductors or Solution conductors

- $\Rightarrow$ Substances which conduct electricity due to movement of ions are called electrolytic conductors or solution conductor.
- $\Rightarrow$ They involve chemical change.
- $\rightarrow$ conductivity increases with rise in temp due to increase ionic mobility.
- $\Rightarrow$ They obey ohm's law as well as Faraday law of electrolysis.

Cell $\rightarrow$ cell is an arrangement which consist electrode and electrolyte.

Electrode $\rightarrow$ It is generally a rod or plate of a good conductor (mostly metal) which is used to supply current in/out of the electrolytic solution.

